

2. GASES:

$T_1 = T_2 = T = \text{Constante} \Rightarrow$ PROCESO ISOTERMO

↓
Ley de Boyle-Mariotte

↓

$$P_1 V_1 = P_2 V_2$$

(1)

$$P_1 = 3'3 \text{ atm}$$

$$V_1 = 7'1 \text{ l}$$

(2)

$$P_2 = 2'9 \text{ atm}$$

$$V_2 = ?$$

Utilizando la ecuación de Boyle-Mariotte tenemos:

$$(3'3 \text{ atm}) \cdot (7'1 \text{ l}) = (2'9 \text{ atm}) \cdot V_2 \Rightarrow$$

$$\Rightarrow V_2 = \frac{(3'3 \text{ atm}) \cdot (7'1 \text{ l})}{(2'9 \text{ atm})} = 8'08 \text{ l} \approx 8'1 \text{ l}$$

$$V_2 = 8'1 \text{ l}$$