

# 1-GASES

$$V = \delta L = \text{Constante} \Rightarrow V_1 = V_2 = V = \delta L$$

(1)

$$T_1 = 23^\circ\text{C} \Rightarrow T_1 = 296\text{ K}$$

$$P_1 = 1015\text{ hPa} \cdot \frac{100\text{ Pa}}{1\text{ hPa}} = 1'015 \cdot 10^5\text{ Pa}, \quad P_2 = ?$$

(2)

$$T_2 = 75^\circ\text{C} \Rightarrow T_2 = 348\text{ K}$$

Proceso a Volumen constante  $\Rightarrow$  Ley Gay-Lussac

$$\boxed{\begin{array}{c} \downarrow \\ \frac{T_1}{P_1} = \frac{T_2}{P_2} \end{array}}$$

$$\frac{296\text{ K}}{1'015 \cdot 10^5\text{ Pa}} = \frac{348\text{ K}}{P_2} \Rightarrow P_2 = 348\text{ K} \cdot \frac{1'015 \cdot 10^5\text{ Pa}}{296\text{ K}}$$

$$P_2 = 1'193 \cdot 10^5\text{ Pa} \approx 1'19 \cdot 10^5\text{ Pa}$$

$$P_2 = 1'19 \cdot 10^5\text{ Pa} = 119000\text{ Pa}$$

$$119000\text{ Pa} \cdot \frac{1\text{ atm}}{101300\text{ Pa}} = 1'18\text{ atm} = P_2$$