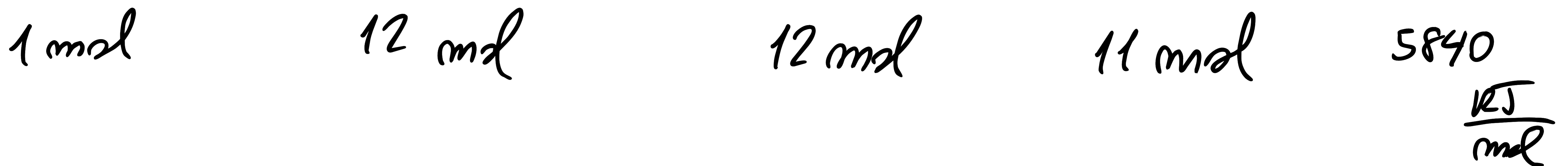
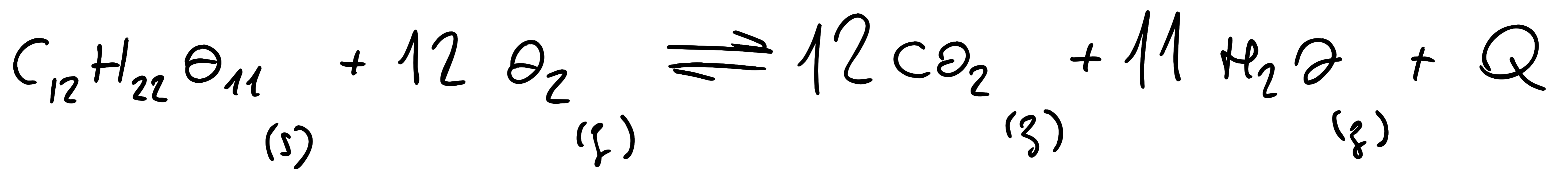


## 23º) EDELVIDES (p. 222)



a) R. Combustión

b)

$$50 \text{ g C}_{12}\text{H}_{22}\text{O}_{11} \cdot \frac{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}}{342 \text{ g C}_{12}\text{H}_{22}\text{O}_{11}} \cdot \frac{11 \text{ mol H}_2\text{O}}{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}} \cdot \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 28'9 \text{ g H}_2\text{O}$$

$$50 \text{ g C}_{12}\text{H}_{22}\text{O}_{11} \cdot \frac{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}}{342 \text{ g C}_{12}\text{H}_{22}\text{O}_{11}} \cdot \frac{12 \text{ mol CO}_2}{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}} \cdot \frac{44 \text{ g CO}_2}{1 \text{ mol CO}_2} = 77'2 \text{ g CO}_2$$

c)  $\Downarrow$

$$0'146 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11} \cdot \frac{12 \text{ mol CO}_2}{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}} \cdot \frac{22'4 \text{ L CO}_2}{1 \text{ mol CO}_2} = 39'3 \text{ L CO}_2$$

d)

$$0'146 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11} \cdot \frac{5840 \text{ kJ}}{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}} = 853'8 \text{ kJ se desprenden al quemar } 50 \text{ g de azúcar.}$$