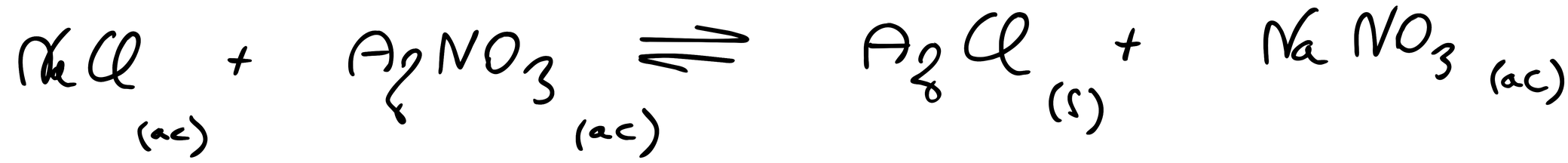
22º) EDELVIVES (p. 272)



$$1'43\,g\ AgCl \cdot \frac{1\,mol\ AgCl}{143'35\,g\ AgCl} \cdot \frac{1\,mol\ NaCl}{1\,mol\ AgCl} = 0'01\,mol\ NaCl$$

$$0'01\,mol_{NaCl} \cdot \frac{1\,L_D}{0'5\,mol\ NaCl} = 0'02\,L_D$$

$$0'02\,L_D \cdot \frac{1000\,ml_D}{1\,L_D} = 20\,ml_D$$